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# Effect of High Pressures on the Formation of Nitric Oxide in Lean, Premixed Flames

*Increasingly stringent regulations are imposed on nitrogen oxides emissions due to their numerous negative impacts on human health and the environment. Accurate, experimentally validated thermochemical models are required for the development of the next generation of combustors. This paper presents a series of experiments performed in lean, premixed, laminar, jet-wall stagnation flames at pressures of 2, 4, 8, and 16 atm. To target postflame temperatures relevant to gas turbine engines, the stoichiometry of the nonpreheated methane–air mixture is adjusted to an equivalence ratio of 0.7. One-dimensional (1D) profiles of temperature and NO mole fraction are measured via laser-induced fluorescence (LIF) thermometry and NO-LIF, respectively, to complement previously published flame speed data (Versailles et al., 2018, "Measurements of the Reactivity of Premixed, Stagnation, Methane-Air Flames at Gas Turbine Relevant Pressures," ASME. J. Eng. Gas Turbines Power, 141(1), p. 011027). The results reveal that, as the pressure increases, the maximum postflame temperature stays relatively stable, and the concentration of NO produced through the flame front remains constant within uncertainty. Seven thermochemical models, selected for their widespread usage or recent date of publication, are validated against the experimental data. While all mechanisms accurately predict the postflame temperature, thanks to consistent thermodynamic parameters, important disagreements are observed in the NO concentration profiles, which highlights the need to carefully select the models used as design tools. The lack of pressure dependence of NO formation that many models fail to capture is numerically investigated via sensitivity and reaction path analyses applied to the solution of flame simulations. The termolecular reaction $H + O_2(+M) \leftrightarrow HO_2(+M)$ is shown to hinder the production of atomic oxygen and to consume hydrogen radicals at higher pressures, which inhibits the formation of nitric oxide through the $N_2O$ pathway.* [DOI: 10.1115/1.4047748]

## Introduction

Nitrogen oxides are involved in the formation of numerous pollutants, such as acid precipitations, ground-level ozone, smog, nutrient contamination of surface water, and fine aerosols that affect human health and the environment [1]. Governments imposed increasingly severe emission standards over the last decades, which induced major modifications to the architecture of combustors. For stationary applications, great improvements were achieved through a better control of the post-flame temperature field, namely, by implementing lean premixed concepts, reducing mixture composition inhomogeneities through improved mixing, and decreasing cooling requirements to maximize the amount of air injected in the combustion chamber. These strategies allowed to minimize NO formation through the thermal (Zel'dovich) route, and achieve sub-10 ppm $NO_x$ emissions [2], a level where the secondary $N_2O$, prompt, and NNH pathways are expected to significantly contribute [3]. An improved understanding of these chemical routes is, therefore, needed to inform the development of the next generation of low-emission combustors. Furthermore, accurate thermochemical models are required since numerical tools are intensively used to reduce development cost, and to design prototypes as close to production as possible. To properly constrain and validate the models, a comprehensive set of independent experimental data, collected in fundamental, laboratory burners minimizing combustor-specific effects [4], must be assembled [5,6].

Several experimental studies, e.g., Refs. [3] and [7–10], investigated nitric oxide formation at low and ambient pressures for a variety of fuels, inert content, and stoichiometries. However, only a few experimental datasets, collected in McKenna [11–14], and counterflow [15] burners, were published at high pressures, and the ability of thermochemical models to accurately predict NO concentrations at these conditions remains unknown. In this paper, the experimental data measured by Versailles et al. [16] in laminar, lean, methane–air, jet-wall stagnation flames, and reported as absolute laser-induced fluorescence (LIF) signal intensities, are remastered using a direct calibration method to yield NO mole fractions. Originally limited to 8 atm, the dataset is extended to 16 atm, and supplemented with one-dimensional (1D) temperature profiles measured via multiline NO-LIF thermometry along the axis of the burner. These measurements, along with the flame speed data provided in Ref. [17], form a complete ensemble of targets against which thermochemical models can be thoroughly validated at postflame temperatures (≈1820 K) and pressures (up to 16 atm) relevant to gas turbine engines. In this paper, the dataset is employed to benchmark the predictions of seven thermochemical models available in the literature [18–25], and the most accurate, published by the National University of Ireland, Galway, is used to investigate the pressure dependence of NO formation in lean premixed flames.

## Experimental Method

The present experiments were performed in the high-pressure, laminar flame facility at McGill University [17]. It consists of a jet-wall, stagnation flame burner installed in an optically accessible pressure vessel (0.29 m inner diameter, 0.86 m long) capable of continuous operation at the full-load pressure of modern aero-derivative gas turbines. The NO concentration and bulk flow temperature were measured via LIF, and NO-LIF thermometry,

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respectively, adapted in the course of the work to high-pressure conditions.

**Jet-Wall, Stagnation Flame Burner.** The jet-wall, stagnation flame burner employed for the current experimental campaign is shown in Fig. 1. A well-mixed stream of methane and air flows through the inner converging nozzle ($D = 10.2$ mm) and stagnates on a water-cooled plate located ≈9 mm downstream of the nozzle assembly. To prevent condensation and surface reactions [26], the plate is maintained at temperatures between 422 K and 489 K, approximately, throughout the experiments. To insulate the flame from the surroundings and improve its stability, it is shrouded by a stream of inert nitrogen gas [27]. In this configuration, the flame stabilizes at the location where its propagation speed matches the axial velocity of the incoming combustible mixture. As shown in Fig. 1, it is compact, lifted, stable, and readily accessible for laser-based diagnostics. For all conditions, the flame position is carefully adjusted to minimize interference from the burner boundaries. The chemical reactivity and production of pollutants are, therefore, sole functions of the properties of the reactant mixture.

During the experiments, the streams of nitrogen, air, and methane are monitored using thermal mass flow controllers calibrated, for each operating pressure, using a dry-piston calibrator. This results in a relative uncertainty on the equivalence ratio of 0.71%. The temperatures of the reactants and stagnation plate are continuously measured using type-K thermocouples connected to a 24-bit isothermal data acquisition card. The pressure ($P$) is regulated by a digital controller connected to a pneumatic valve. The response of the diaphragm-based, piezoresistive sensor is calibrated against a NIST traceable gauge, yielding an uncertainty on $P$ of 0.03 atm (3040 Pa). Further details on the burner, vessel, and equipment are found in Ref. [17].

**NO Laser-Induced Fluorescence.** The concentration of nitric oxide is measured via two-dimensional, laser-induced fluorescence in the linear (weak irradiation) regime. A wavelength-tunable dye laser (Sirah Lasertechnik GmbH, Grevenbroich, Germany) filled with a solution of Coumarin 450 dye is pumped by the third harmonic of a Nd:YAG laser (Spectra-Physics Quanta-Ray Pro-230, Santa Clara, CA) to produce a beam at a wavelength of ≈452 nm. Its passage through a beta barium borate doubling crystal yields deep ultraviolet light at ≈226 nm. To obtain a fairly homogeneous energy profile, the beam travels through a homogenizer made of two planoconvex, cylindrical microlens arrays that create a number of smaller laser beams in the vertical direction, which are overlapped on the centerline axis of the burner by a planoconvex cylindrical lens to form an ≈9 mm tall, ≈1 mm wide, laser sheet [28].

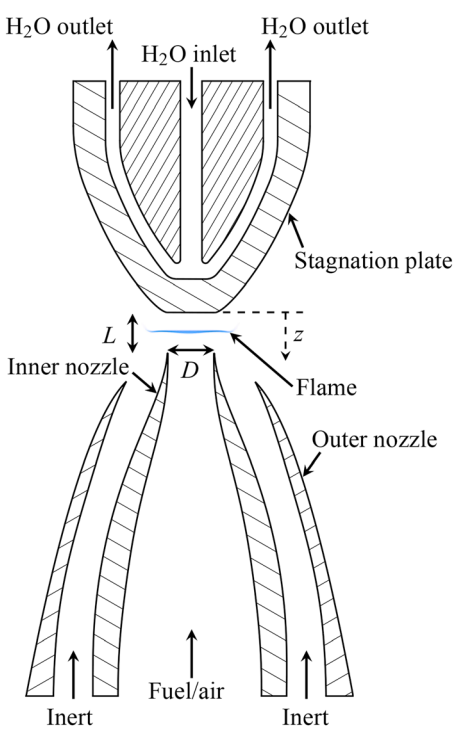


**Fig. 1 Jet-wall stagnation burner [1]**

Bessler et al. [29] studied many excitation and detection schemes, and concluded that A-X(0,0) excitation followed by red-shifted detection offers the highest NO laser-induced fluorescence signals, while minimizing interfering $O_2$-LIF. Consequently, the laser wavelength is adjusted to the spectral feature at ≈226.03 nm recommended by DiRosa et al. [30], which comprises the $S_{21}(4.5)$, $P_1(23.5)$, $P_{21} + Q_1$ (14.5), and $Q_2 + R_{12}$ (20.5) transitions of the A-X(0,0) electronic system of NO. The spontaneously emitted photons pass through a long-pass filter (Semrock LP02-224R (IDEX Health & Science, LLC Rochester, NY)), which blocks most of the Rayleigh scattering and transmits the A-X (0,1) and subsequent vibrational emission bands of nitric oxide, before being collected by an achromatic UV lens (Sodern Cerco 2178 (Sodern Limeil-Brévannes, France)) mounted on extension rings for improved magnification. The photons are recorded by an intensified CCD camera (Dicam Pro, Gen II S-20 photocathode (PCO AG Kelheim, Germany)) exposed for 300 ns, with a 4 (vertical) × 8 (horizontal) pixel binning to increase the signal-to-noise ratio.

By invoking laser-induced fluorescence in the linear (weak irradiation) regime, the NO-LIF signal per unit laser energy for a given pixel on the ICCD detector, $F_{NO}$, can be written as [16]:

$$F_{NO} = \frac{S_{NO}}{E_L} = \frac{S_{Raw} - S_{Bkg_0}}{E_L} - F_{Bkg_1} \tag{1}$$

where $S_{NO}$ is the NO-LIF signal; $E_L$ is the mean of the time-integrated laser pulse power profiles measured with a high-speed photodiode (Thorlabs DET10A (Thorlabs Inc., Newton, NJ)) and an 1 GHz sampling rate oscilloscope (Picoscope 2207B (Pico Technology Neots, Cambridgeshire, UK)); $S_{Raw}$ is the raw signal measured by the camera (average of 5000 images with laser irradiation); $S_{Bkg_0}$ is a combination of laser-independent background signals, e.g., flame chemiluminescence, camera dark noise, surface reflections, etc. (measured by averaging 2000 images of the flame without laser excitation); and $F_{Bkg_1}$ is the signal per unit energy of laser-dependent interfering signals, such as $O_2$-LIF, $CO_2$-LIF, and scattering [29]. For each location along the axis of the burner, the $F_{NO}$ values of 20 pixels (≈4.2 mm) in the vicinity of the centerline are averaged to yield a 1D profile of LIF signal intensity, which is then corrected for spatial variations in the laser sheet energy using the average of 2000 LIF images measured in a cold, nonreacting stream of 300 ppm of NO in $N_2$.

To determine $F_{Bkg_1}$, the method of Versailles et al. [16], which does not assume that the background intensity is invariant to the laser wavelength as typically done [15,31], is employed. As shown in Fig. 2 and detailed in Ref. [16], two excitation spectra without (orange curve) and with 53.7 ppm of nitric oxide seeded in the reactants (green curve) are measured at laser wavelengths ranging from 226.005 nm to 226.115 nm. One hundred twenty images are averaged at each wavelength, corrected for $S_{Bkg_0}$, and normalized by $E_L$. The subtraction of both spectra corresponds to the excitation spectrum of pure NO (blue curve). An iterative algorithm then determines the appropriate fraction ($C_{Bkg}$) of the NO spectrum (blue curve) to remove from the seeded trace to yield a background signal devoid of NO resonant features (black curve). The procedure is repeated for each axial location to obtain a one-dimensional profile of $C_{Bkg}$, which is smoothed by a least-squares adjusted sextic Bézier curve. To further reduce random oscillations in the axial noise profile, 5000 images in seeded and unseeded flames are averaged at the on-resonance wavelength.



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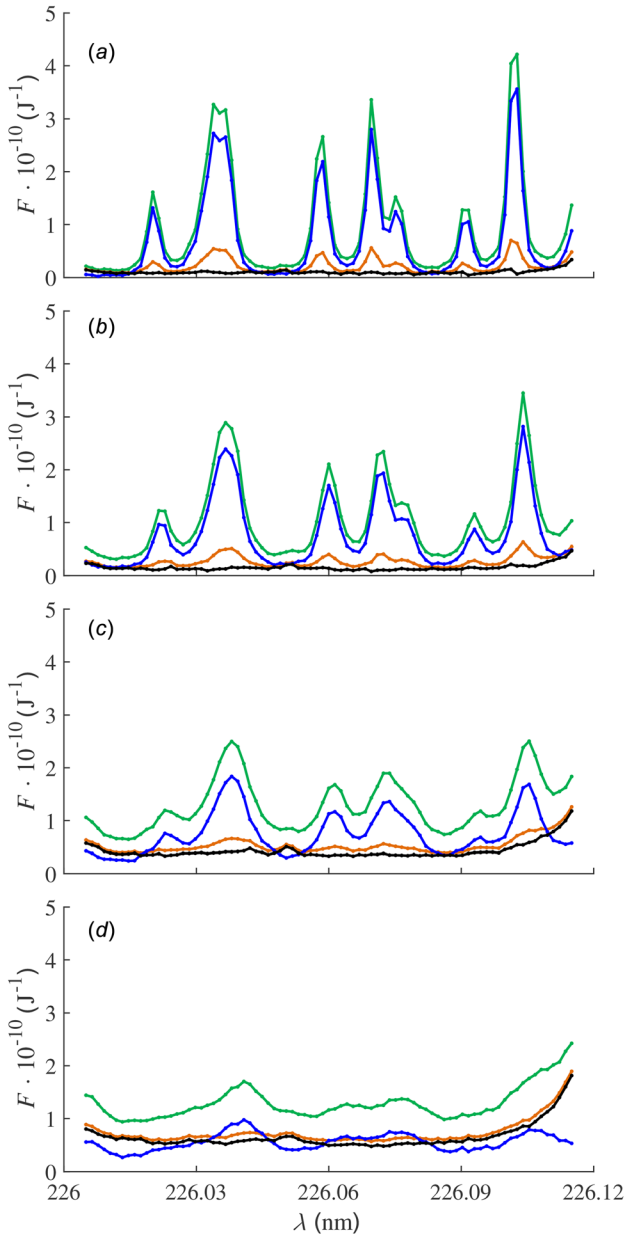


**Fig. 2 Iterative background determination method applied in the postflame region of the 2 atm (*a*), 4 atm (*b*), 8 atm (*c*), and 16 atm (*d*) flames. Legend: excitation spectra with (green) and without (orange) NO seeding, excitation spectra of pure NO (blue), and estimated background signal (black). On-resonance excitation is performed on the absorption band located at ≈226.04 nm. (For interpretation of the references to color, the reader is referred to the online version of this paper.)**

For each axial position, the difference between both signals is multiplied by the local value of the sextic polynomial, and subtracted from the fluorescence intensity in the seeded flame, to yield the 1D profile of $F_{\mathrm{Bkg}_1}$ to be inserted in Eq. (1).

Figure 2 illustrates the inherent difficulty in performing NO-LIF measurements at supra-atmospheric conditions; as $P$ increases, the fluorescence signal of NO (blue curve) diminishes,[2] while the intensity of the background interference rises. Namely, at 16 atm, the significantly broadened NO spectral features are barely recognizable in the unseeded spectrum. Under such conditions, the background signal determination procedure described above is ineffective. An alternate technique, inspired from Refs. [15] and [31], is then implemented. A priori, on- and off-resonance wavelengths where the interfering signal is equal are unknown. However, taking advantage of the data collected at 8 atm and processed with the method of Versailles et al. [16], a range of wavelengths from 226.06 nm to 226.08 nm is identified over which the background trace is relatively constant, while the NO-LIF signal varies. The spectrum of the background signal is then obtained by iteratively determining the fraction of the NO-LIF signal (blue curve in Fig. 2(*d*)) to remove from the seeded trace (green curve) to minimize the sum of squares of the difference between the background intensity at each of the wavelengths in the range 226.06–226.08 nm. Because of the wide variability in the $C_{\mathrm{Bkg}}$ values in the postflame region, the adjustment of the sextic Bézier polynomial is unsuccessful. Consequently, the value of $C_{\mathrm{Bkg}}$ is calculated as the average of 10 pixels neighboring a single position approximately halfway between the flame front and the stagnation surface, and used to determine $F_{\mathrm{Bkg}_1}$ as previously described.

Laser-induced fluorescence in the linear regime can be described by Eq. (2), where $X_{\mathrm{NO}}$ is the mole fraction of NO molecules; $N(T,P)$ is the number density; $f_{\mathrm{LIF}}$, generally provided by a LIF model, is the number of photons emitted per unit molecule of NO, and unit laser energy; $f_{\mathrm{B}}(T)$ is the Boltzmann fraction of NO molecules in the state excited by the laser; $\lambda$ is the laser wavelength; $\Delta\nu_{\mathrm{L}}$ is the spectral width of the laser; $\Gamma(\Delta\nu_{\mathrm{L}},T,P,X_i)$ is the dimensionless overlap fraction; $B_{12}$ is the Einstein constant of photon absorption; $A_{21}$ is the rate constant of spontaneous emission; $Q_{21}(T,P,X_i)$ is the rate constant of nonradiative collisional quenching; and $C_{\mathrm{opt}}$ is the optical constant of the LIF system.

$$F_{\mathrm{NO}} = X_{\mathrm{NO}} \cdot N \cdot f_{\mathrm{LIF}}(f_{\mathrm{B}}, \lambda, \Delta\nu_{\mathrm{L}}, \Gamma, B_{12}, A_{21}, Q_{21}) \cdot C_{\mathrm{opt}} \tag{2}$$

Some of these parameters can be calculated. However, species and temperature profiles must be hypothesized, which introduces uncertainty. Instead, a direct calibration procedure is used here. Five thousand LIF images collected in a flame seeded with 53.7 ppm of NO are averaged to yield $F_{\mathrm{NO,s}}$ in Eq. (3). To remove the contribution of the nitric oxide produced by the flame, the LIF signal measured in the unseeded flame, $F_{\mathrm{NO}}$, is deducted of $F_{\mathrm{NO,s}}$. Assuming that the seeded NO molecules are inert through the flame, a condition verified numerically and generally valid for lean flames [15], the net LIF signal can be written as:

$$F_{\mathrm{NO,s}} - F_{\mathrm{NO}} = 53.7\,\mathrm{ppm} \cdot N \cdot f_{\mathrm{LIF}} \cdot C_{\mathrm{opt}} \tag{3}$$

Given the relatively small amount of nitric oxide injected in the reactant stream, it is assumed that the bulk temperature and concentration of the main combustion products are identical in both, seeded and unseeded, flames. Hence, all parameters other than the NO mole fraction in the right-hand-side term of Eqs. (2) and (3) are equal. It follows that the experimental mole fraction of NO molecules is calculated as:

$$X_{\mathrm{NO}}\,[\mathrm{ppm}] = \frac{F_{\mathrm{NO}}}{F_{\mathrm{NO,s}} - F_{\mathrm{NO}}} \cdot 53.7\,\mathrm{ppm} \tag{4}$$

To develop Eq. (4), it is hypothesized that $C_{\mathrm{opt}}$ is identical in Eqs. (2) and (3). However, self-absorption (trapping) of the fluorescence signal by the surrounding nitric oxide molecules, which are at different concentrations in the seeded and unseeded flames, could invalidate this assumption. To assess the importance of this phenomenon, photon mean-free paths of 5.9 m (at 16 atm) to 15.5 m (at 2 atm) were calculated.[3] These values are 2–3 orders of magnitude larger than the characteristic radius of the flame

[2]For lean flames, the consumption of nitric oxide seeded into the reactants is negligible. Therefore, the blue spectra in Figs. 2(*a*)–2(*d*) correspond to a concentration of ≈53.7 ppm of NO, which does not vary with $P$.

[3]The calculations were performed for the seeded (calibration) flames. At these conditions, the concentration of NO molecules is expected to be the highest and, therefore, self-absorption most critical.



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**Table 1 Experimentally determined boundary conditions**

| $P$ (atm) | $\phi$ | $l$ (mm) | $u_{\text{inlet}}$ (m/s) | $du/dz\|_{\text{inlet}}$ (1/s) | $T_{\text{inlet}}$ (K) | $T_{\text{wall}}$ (K) |
|---|---|---|---|---|---|---|
| 2 (0.03) | 0.7 (0.005) | 8.17 (0.05) | 0.2029 (0.0008) | 48.33 (4.91) | 291 (2) | 422 (5) |
| 4 (0.03) | 0.7 (0.005) | 8.12 (0.05) | 0.1288 (0.0004) | 28.43 (2.31) | 293 (2) | 445 (5) |
| 8 (0.03) | 0.7 (0.005) | 6.62 (0.05) | 0.0754 (0.0002) | 15.10 (0.94) | 292 (2) | 440 (5) |
| 16 (0.03) | 0.7 (0.005) | 5.15 (0.05) | 0.0473 (0.0001) | 10.56 (0.75) | 293 (2) | 489 (5) |

Absolute uncertainties are shown within parentheses.

(approximately equal to $D$). Therefore, self-absorption of the NO-LIF signal is negligible. Also, it must be noted that absorption of the LIF signal by main combustion products, such as hot carbon dioxide and water vapor [29], if significant, does not affect the calibration since their concentration is the same in the seeded and unseeded flames.

**NO-LIF Thermometry.** The temperature of the gas mixture is measured using a multiline, one-dimensional, NO-LIF thermometry method [32,33] using the equipment described in NO Laser-Induced Fluorescence section. The reactant stream is seeded with 537.4 ppm of NO, a concentration expected to have a minimal impact on the flame properties, namely, the temperature of the products [33]. The wavelength of the laser is varied to cover multiple features of the NO excitation spectrum, whose respective intensity is temperature-dependent as shown in Fig. 3. For each axial position, a fitting procedure adjusts a series of parameters (the relative intensity of the spectrum (multiplier), the baseline (background) signal and its inclination, the laser line width (Gaussian and Lorentzian contributions), and the wavelength shift of the whole spectrum) to minimize the sum of squares of the difference between the measured and simulated excitation spectra, and extract the bulk gas temperature to an accuracy of ≈5% at 2000 K [33].

Traditionally [8,32,33], the three-level laser-induced fluorescence model LIFSim [34] is employed to generate the theoretical excitation spectra required by the fitting procedure. However, in the course of the work, it was observed that some features of the experimental spectra did not exactly overlap the simulated traces. A detailed comparison of the spectral position of the different transitions in the A-X (0,0) system of NO revealed significant differences between the predictions of LIFSim and LIFBASE [35]. A different methodology is, therefore, implemented here to generate the theoretical traces. A database of excitation spectra, solely accounting for Doppler broadening, is assembled for a wide range of temperatures based on the predictions of LIFBASE. The same parameters listed above are adjusted by the fitting procedure, the only difference being that the width of the Lorentzian distribution, which now needs to be convolved with the spectra of the database since this step was previously performed by LIFSim, is the combination of the laser spectral profile and the pressure-dependent collisional broadening. Owing to the more accurate spectral positions in LIFBASE, a more than threefold reduction in the residuals of the least-squares method is observed for the 2 atm flame.

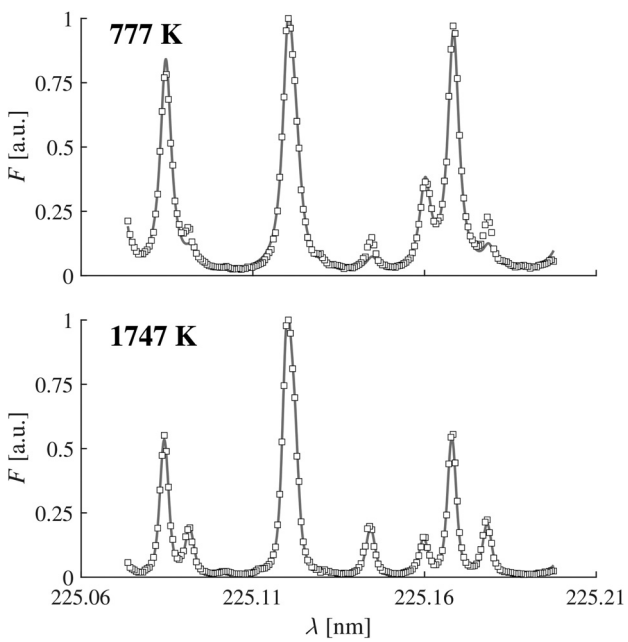


**Fig. 3 Excitation spectra obtained within the front (top), and in the postflame region (bottom) of the 2 atm flame. Legend: experimental data (square symbols), and numerical spectra least-squares adjusted to the experimental data (gray curves).**

As shown in Fig. 2, the background interference becomes increasingly important in comparison to the NO fluorescence signal at higher pressures. If improperly quantified, the interfering signals can have a detrimental impact on the accuracy of the measurements [33]. Consequently, a second excitation spectrum in an unseeded flame is subtracted of the trace collected in the seeded flame. This removes the contribution of the interfering signals, and yields the excitation spectrum of pure NO (analogous to the blue curves in Fig. 2).

For all pressures, excitation scans (with and without NO seeding) with a resolution of 0.623 pm are performed for wavelengths extending from 225.0736 nm to 225.1976 nm. Fifty images are averaged at each wavelength, corrected for $S_{\text{Bkg}_0}$, normalized by the laser energy ($E_{\text{L}}$) and, finally, processed into spectra. While this range of laser wavelengths offers a favorable sensitivity of the spectra to the gas temperature, it suffers from weak NO-LIF signal intensities at high pressures, namely, in the colder region through and upstream of the flame. To obtain more reliable data in this zone, a second set of measurements averaging 66 fluorescence images collected at 150 wavelengths (2.0134 pm resolution) between 226.2145 nm and 226.5045 nm is performed for the 8 and 16 atm flames.

**Quasi-One-Dimensional Flame Modeling.** The jet-wall burner was purposely designed to yield flat, axisymmetrical flames [36,37] producing 1D fields of species concentration, temperature, density, and axial velocity in the vicinity of the centerline axis. Coupled with an axisymmetrical stream function, they allow to reduce the three-dimensional continuity, Navier–Stokes, and energy and species conservation equations to a quasi-1D formulation [38], implemented in Cantera 2.3 [39]. Radiative heat loss and the mixture-averaged formulation of the transport coefficients are included in the flame simulations, converged to absolute and relative tolerance levels of $10^{-9}$ and $10^{-5}$, respectively, on highly resolved meshes achieving gradient and curvature parameters of 0.05 and 0.075, respectively. The boundary conditions (BC) for the flame simulations are reported in Table 1. The equivalence ratio is obtained from the mass flow rates of fuel and air provided by the thermal mass flow controllers, the temperatures collected using thermocouples, and the velocity BCs measured by particle tracking velocimetry [17,37].

The flame simulations are performed with seven thermochemical models selected for their wide usage in industry and academia,



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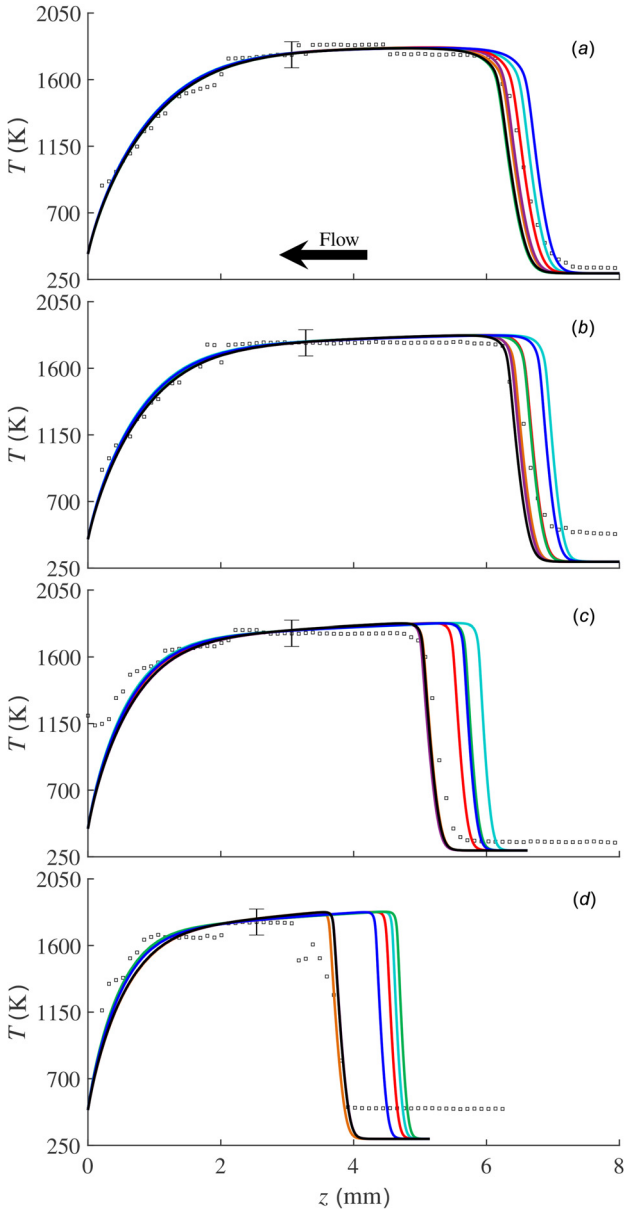


**Fig. 4 One-dimensional profiles of temperature at 2 atm (*a*), 4 atm (*b*), 8 atm (*c*), and 16 atm (*d*). The plate is located at $z = 0$ mm, and the outlet of the nozzle at $z \approx 9$ mm. Legend: experiments (symbols), NUIG (orange), CRECK (purple), CSE (cyan), KON (green), SD (blue), GRI (red), and ARG (black).**

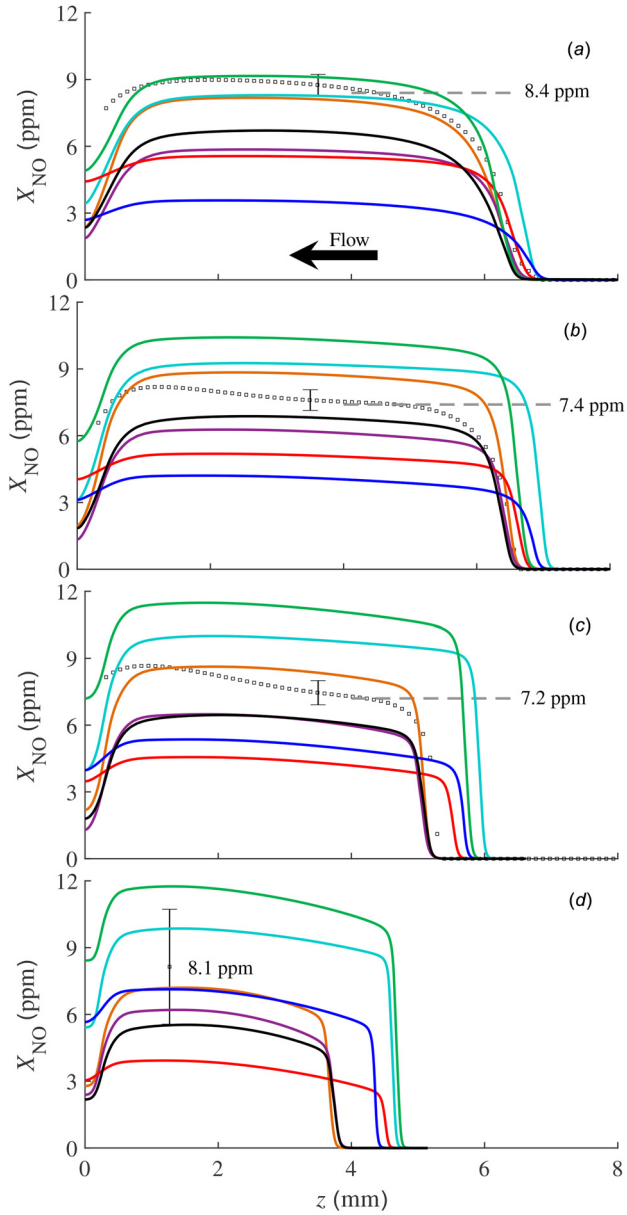


**Fig. 5 One-dimensional profiles of NO mole fraction at 2 atm (*a*), 4 atm (*b*), 8 atm (*c*), and 16 atm (*d*). Same legend as Fig. 4. The error bars account for the accuracy of the LIF method, as well as the effect of the uncertainty of the BCs estimated with flame simulations performed with the NUIG model.**

detailed modeling of the chemistry, date of release, or design philosophy: GRI-Mech 3.0 (GRI) [18]; the 2016 version of the San Diego mechanism (SD) with the nitrogen chemistry published in 2004[4] [19]; the Combustion Science & Engineering, Inc. (CSE) model [20]; the high-temperature version of the mechanism from the National University of Ireland, Galway (NUIG) [21,22]; the Konnov mechanism release 0.6 [23]; the high-temperature chemistry (version 1412) from the CRECK Modeling Group (CRECK) at Politecnico di Milano [24]; and the mechanism provided with the recently published review paper on the nitrogen chemistry of Glarborg et al. [25] (ARG).

[4]An updated version of the NO chemistry was released in 2018. However, it does not include all NO formation routes known to proceed in flames of gaseous hydrocarbon fuels. Namely, the prompt (Fenimore) pathway is excluded. For this reason, the more complete version 2004 is used here.

## Experimental and Numerical Results

The experimental temperature profiles along the axis of the burner are shown in Fig. 4 (symbols), where the water-cooled stagnation plate is located at $z = 0$ mm, and the outlet of the nozzle at $z \approx 9$ mm. As expected, the temperature is initially low, then rises rapidly through the front to reach a value close to the adiabatic flame temperature, and finally decreases in the thermal boundary layer (TBL) close to the plate due to convective heat loss. The temperature in the postflame region outside of the TBL is observed to minimally change with the pressure. This indicates that dissociation is minimal at such low flame temperatures. Otherwise, increasing the pressure would shift the equilibrium to larger, more stable molecules as per the Le Chatelier's principle, and a rise in flame temperature would be observed. It is noted, principally at 4 and 16 atm, that the temperature of the reactants is significantly higher than the value measured by the thermocouples (see Table 1). In this region, the residuals of the least-squares



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fitting routine are much higher (about one order of magnitude) than in the postflame region, which is expected to negatively impact the accuracy of the LIF thermometry.

The predictions of the seven thermochemical mechanisms are shown by the curves in Fig. 4. All models are consistent with each other and with the experimental data in the postflame region, which indicates that the models include an accurate, consistent set of thermodynamic coefficients. The slight reduction in the predicted temperature upstream of the TBL, which is caused by radiative heat loss, is not observed in the experimental data. This discrepancy is, however, within the bounds of uncertainty. Also, the temperature reduction in the thermal boundary layer is well captured by all models. The principal differences between the models and against the experimental data are observed in the flame front region. The spread in flame position is due to inconsistent predictions of flame reactivity [17]. Namely, in stagnation flows, an overestimation of the flame speed will result in a front located upstream of the experimental flame, and vice-versa.

The experimental profiles of NO mole fraction are shown in Fig. 5. As expected, nitric oxide is rapidly formed through the front mostly via the $N_2O$ pathway, and then relatively slowly in the postflame region, principally via the thermal route [16]. In the thermal boundary layer, $X_{NO}$ decreases indicating that a fraction of the nitric oxide molecules formed through the flame are consumed. As explained above, the measurements at 16 atm are plagued by a significant amount of noise, such that only a single data point could be extracted. However, the general shape of the mole fraction profile should be similar to lower pressures as suggested by the simulations. As concluded in Ref. [16] based on the nominal concentrations at 2, 4, and 8 atm, the amount of NO formed through the flame front (numerical values indicated on the plots) slightly decreases monotonically with the pressure. However, considering the width of the newly added error bars, as well as the new data at 16 atm, a more appropriate conclusion is that NO formation in lean, premixed flames, dominated by the $N_2O$ pathway, does not present a significant dependence to the pressure. This has important implications for engines. Namely, the current dataset collected at pressures and postflame temperatures relevant to gas turbine combustors indicates that single-digit NO emissions are achievable from low to high loads, if the premixer is carefully designed to provide very low unmixedness.

The NO concentration predictions of the seven thermochemical models are shown by the curves in Fig. 5. A wide variability is observed for all investigated pressures. Furthermore, many models provide $X_{NO}$ values that significantly differ from the experimental data, i.e., beyond the uncertainty bounds, which demonstrates the relevancy of the current dataset as optimization and validation targets for model development. Given the consistent temperature predictions presented in Fig. 4, it follows that these discrepancies are not caused by thermal effects, but rather by an incorrect modeling of the chemistry. In particular, the GRI model, which is widely used in industry and academia, under-predicts the experimental data for all conditions (as much as 50% at 8 atm). This implies that a concept, developed using the GRI mechanism to achieve a given emission level, could well miss the target once built. Also, many models do not accurately capture the pressure dependence of the NO mole fraction. Namely, the CSE, SD, and KON mechanisms predict a continuous increase in $X_{NO}$ with the pressure, and these variations are significant in view of the experimental uncertainty. Namely, the values provided by the KON mechanism are within the bounds of error at 2 atm, but completely outside at 16 atm, while the opposite is observed for the SD model. Finally, although they sometimes provide numerical values outside of the experimental bounds of uncertainty, the most recently published CRECK, ARG, and NUIG models generally present the best performance in terms of pressure dependence and NO mole fraction, as well as flame speed and location as demonstrated in Ref. [17]. These observations highlight the importance of properly selecting the chemical models used in the design of low-emission engines.

## Discussion

Figure 6 presents a reaction path analysis (RPA) applied to the solution of the 8 atm flame simulated with the NUIG mechanism, which is shown in Figs. 4 and 5 to provide the most accurate predictions in terms of flame location and NO mole fraction. A cylindrical control volume extending from the inlet of the computational domain to a position 2 ms ($\tau = 2$ ms) downstream of the maximum CH concentration (flame marker) is applied to the flame solution, and the flux of nitrogen atoms (a conserved scalar) transferred between every species as the reactants are transformed into products is calculated with the formulation presented in Refs. [37] and [40] inspired by Grcar et al. [41]. The arrows are colored according to the logarithmic sensitivity of the NO mole fraction at $\tau = 2$ ms to the specific rate ($k_j$) of the reactions making the individual channels (L.S.$(X_{NO,\tau=2\,ms}, k_j)$). The brute-force sensitivity analysis was repeated for each pressure and all 1847 reactions of the NUIG model to produce Fig. 7, which reports the logarithmic sensitivity (L.S.) values of the ten reactions that most impact NO formation (reactions ranked according to the sum of squares of the L.S. values at the four pressures). For all pressures, the absolute value of the logarithmic sensitivity decreases rapidly toward the right of the plot, which indicates that only a few, common reactions determine the formation of NO. Namely, reactions belonging to the $N_2O$ and thermal (Zel'dovich) pathways are observed. Reactions from the hydrogen submodel are also included, many of them known to have a significant influence on the overall flame reactivity. This clearly demonstrates the need for an accurate base chemistry in order to accurately predict NO formation.

Figure 6 shows that the vast majority of the nitrogen atoms originally transported by $N_2$ enters the $N_2O$ route initiated by the termolecular reaction:

$$N_2 + O(+M) \leftrightarrow N_2O(+M) \tag{5}$$

The other, prompt (Fenimore), NNH, and thermal, routes play a minor role through the front of the current lean, low-temperature flames. Once formed, the $N_2O$ molecules are either transformed back into $N_2$ by the reverse of reaction (5) and by reaction (6); transformed into NO via reactions (7) and (8); or simply escape the control volume (gray-dashed arrow in the RPA diagram).

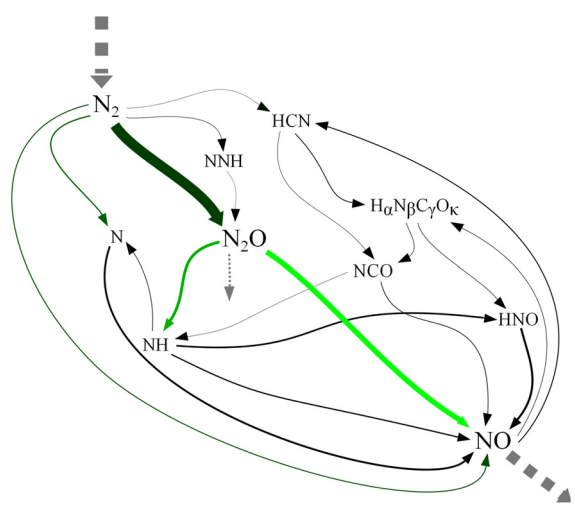


**Fig. 6 Reaction path analysis tracking the flux of nitrogen atoms in the 8 atm flame. The inlets of the RPA control volume and computational domain coincide, and the outlet of the CV is placed at $\tau = 2$ ms. The dashed gray arrows indicate fluxes of N across the boundaries of the control volume. The other arrows are colored according to a linear black (L.S.$(X_{NO,\tau=2\,ms}, k_j) = 0$) to green (L.S.$(X_{NO,\tau=2\,ms}, k_j) > 0$) color map.**



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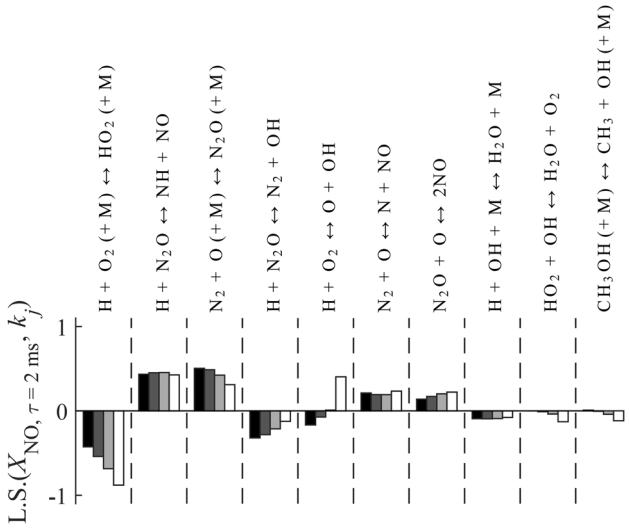


**Fig. 7 Logarithmic sensitivity of the nitric oxide mole fraction at an axial position 2 ms downstream of the flame front. The bands are colored based on a black (2 atm) to white (16 atm) colormap.**

$$H + N_2O \leftrightarrow N_2 + OH \tag{6}$$

$$N_2O + O \leftrightarrow NO + NO \tag{7}$$

$$N_2O + H \leftrightarrow NO + NH \tag{8}$$

The forward rate of reaction (5) can be written as:

$$\begin{aligned} q_F &= [N_2] \cdot [O] \cdot [M] \cdot k_F \\ &\approx X_{N_2} \cdot X_O \cdot \left(\frac{P}{RT}\right)^3 \cdot k_F \end{aligned} \tag{9}$$

where $P$ is the pressure, $R$ is the universal gas constant, $T$ is the temperature, and $k_F$ is the specific rate of the forward reaction. Based on the cubic term in Eq. (9), the pressure could be expected to significantly raise $q_F$ and, consequently, the mole fractions of $N_2O$ and NO. However, as shown in Fig. 5, $P$ has a very weak influence on the formation of NO, even though the pressure is multiplied by a factor of 8 in the current experiments.

Figure 8(*a*) presents the forward rate of reaction (5). It is noted that the maximum value of $q_F$ increases with the pressure, however, not cubically; the maximum rate rises by a factor of ≈29, while it would be expected to be multiplied by 512 ($8^3$) based on Eq. (9) (in fact, more than 512 when considering the pressure dependence of $k_F$). Furthermore, the backward reaction accelerates rapidly as $N_2O$ molecules are created, such that significant values of $q_{Net}$ are limited to a thin layer within the flame front region (Fig. 8(*b*)). Reaction (6), which proceeds preferentially in the forward direction, was observed to consume $N_2O$ molecules at a much slower rate than the reverse of reaction (5). Furthermore, as shown in Fig. 8(*c*), the net rates of reactions (7) and (8) are negligible in comparison to the backward rate of reaction (5) (Fig. 8(*a*)). It follows that the mole fraction profiles of $N_2O$, shown in Fig. 9(*b*), are primarily determined by the net rate of reaction (5). Namely, for all pressures, $X_{N_2O}$ rises rapidly through the front, reaches a maximum value that moderately varies with the pressure, and then decays in the postflame region.

The mole fraction of $N_2$ is mainly dictated by the composition of the reactants, and does not change significantly with the pressure. Therefore, referring to Eq. (9), the lower than expected rise in the forward rate of reaction (5) is related to the mole fraction of atomic oxygen. As shown in Fig. 9(*a*), $X_O$, namely, its maximum value achieved in the flame front region, decreases significantly with the pressure. To identify the main reactions causing this reduction, the logarithmic sensitivity of the peak value of $X_O$ to the specific rate of the reactions was calculated at each pressure (see Fig. 10). As for nitric oxide, only a very small subset of the chemistry model has a significant influence on the concentration of atomic oxygen. Specifically, the termolecular reaction (10), presenting large negative values of L.S. is the most dominant. This reaction, well-known for its chain-terminating effect leading to the second explosion limit in hydrogen-oxygen mixtures [42], competes with the chain-branching reaction $H + O_2 \leftrightarrow O + OH$ that forms oxygen radicals. As the pressure increases, the relative importance of the termolecular reaction grows, which hinders the formation of atomic oxygen. As a consequence, the forward rate of reaction (5) is reduced, and the formation of $N_2O$ molecules inhibited.

$$H + O_2(+M) \leftrightarrow HO_2(+M) \tag{10}$$

The net rates of reactions (7) and (8), which primarily evolve in the forward direction, are presented in Fig. 8(*c*). Interestingly, significant values of $q_{Net}$ for these reactions are only observed within the flame front region, where the radicals required for their progress are available. That is, analogous to the prompt (Fenimore) route, the $N_2O$ pathway is a rapid mechanism of NO formation that proceeds within the front. Similarly to Eq. (9), applying the law of mass action suggests that the rate of these reactions should increase by a factor of 64 between 2 and 16 atm, if the mole fractions of O, H, and $N_2O$, were to remain constant. However, the maximum value of $q_{Net}$ only rises by factors of ≈20 and ≈12 for reactions (7) and (8), respectively. As discussed above, the termolecular reaction (10) has the strongest negative impact on the formation of oxygen atoms. Hence, its influence is not limited to the formation of $N_2O$

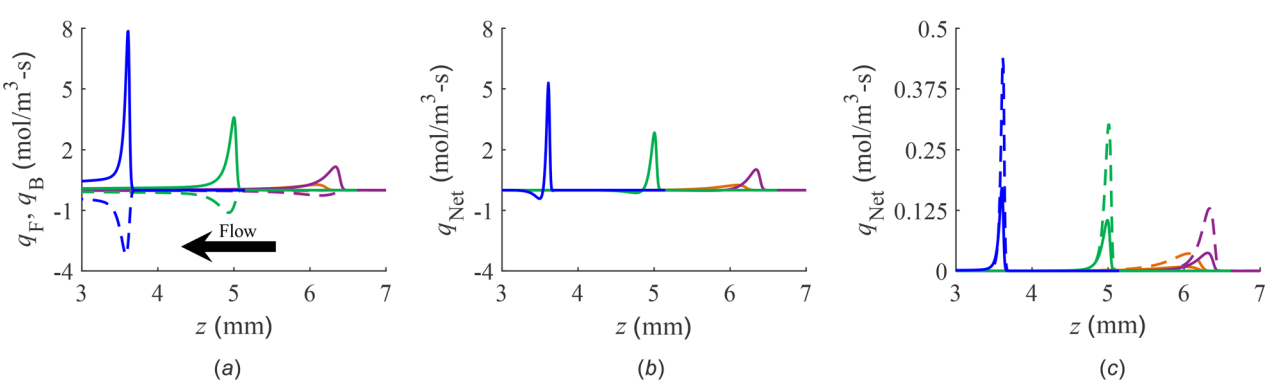


**Fig. 8 Profiles of forward (solid curves, plot a), backward (dashed curves, plot a), and net (*b*) reaction rates of the reaction $N_2$+O(+M) ↔ $N_2$O(+M), and net reaction rates of reactions $N_2$O+H ↔ NO+NH (dashed curves, plot c) and $N_2$O+O ↔ NO+NO (solid curves, plot c). Legend: $P$ = 2 atm (orange), $P$ = 4 atm (purple), $P$ = 8 atm (green), and $P$ = 16 atm (blue).**



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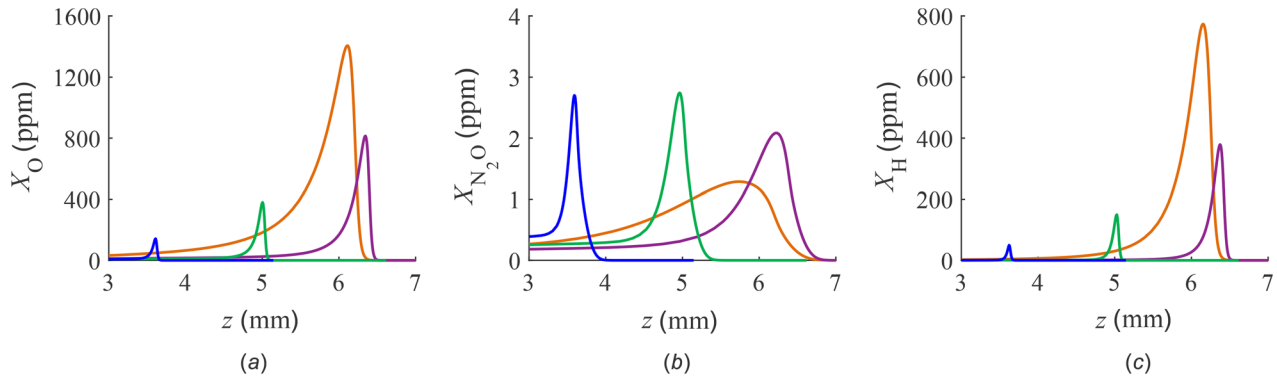


**Fig. 9 Mole fraction profiles of atomic oxygen (*a*), $N_2O$ (*b*), and atomic hydrogen (*c*). Same legend as Fig. 8.**

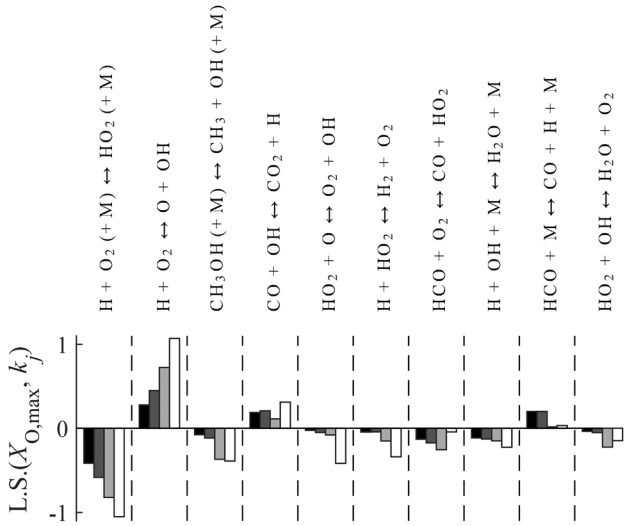


**Fig. 10 Logarithmic sensitivity of the maximum atomic oxygen mole fraction to the specific rate of the individual reactions. Same legend as Fig. 7**

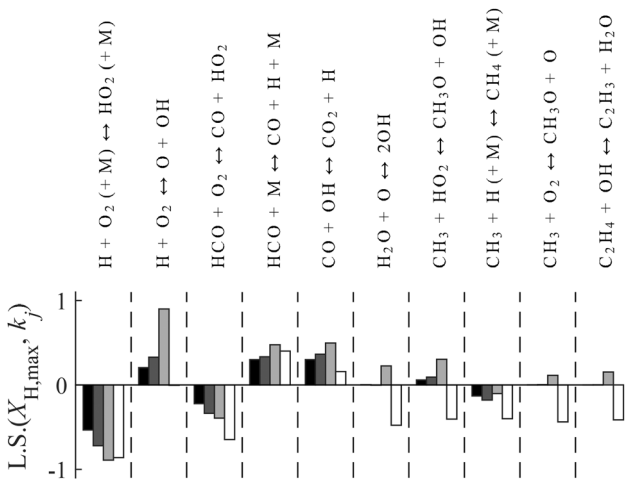


**Fig. 11 Logarithmic sensitivity of the maximum atomic hydrogen mole fraction to the specific rate of the individual reactions. Same legend as Fig. 7.**

molecules; it also inhibits their transformation into nitric oxide via reaction (7). Figure 11 shows the logarithmic sensitivity of the maximum value of $X_H$ to the specific rate of the reactions. Again, the termolecular reaction (10), which consumes the hydrogen radical, is the most dominant. By hampering the formation of atomic hydrogen (see Fig. 9(*c*)), it reduces the rate of reaction (8), and inhibits the formation of NO at higher pressures.

## Conclusion

This paper presents one-dimensional profiles of nitric oxide mole fraction and temperature measured using laser-induced fluorescence, and multiline, NO-LIF thermometry, respectively. The experiments were performed in a jet-wall, stagnation burner at pressures of 2, 4, 8, and 16 atm. The nonpreheated methane–air mixture was adjusted to an equivalence ratio of 0.7 to achieve postflame temperatures of ≈1820 K. For these conditions, nitric oxide is rapidly formed through the front via the $N_2O$ pathway, and then relatively slowly in the postflame zone by the thermal (Zel'dovich) route. While NO formation could be expected to be significantly favored at high pressures due to the termolecular nature of the initiation reaction $N_2 + O(+M) \leftrightarrow N_2O(+M)$, the experimental data instead reveal that, within the bounds of uncertainty, the pressure has a very weak impact on the NO mole fraction, which remains below 10 ppm. This demonstrates that single digit NO emissions are achievable in dry low emission combustors, if the postflame temperature field is properly managed to hinder the thermal (Zel'dovich) route.

The experimental data were used to validate the predictions of seven thermochemical models. While significant differences in the flame reactivity caused discrepancies in the flame front position, all models provide postflame temperatures consistent with the experimental data. On the other hand, a significant variability is observed among the numerical values of NO mole fraction provided by the models, and against the experiments. Namely, the GRI-Mech 3.0 mechanism, widely used in academia and industry, is shown to consistently under-predict the experimental data. Furthermore, many thermochemical models present an incorrect pressure dependence of NO mole fraction. This highlights the need to properly select the models employed for the design of combustion systems.

A detailed study of the nitrogen chemistry, applying sensitivity and reaction path analyses to the flame solutions obtained with the model of the National University of Ireland, Galway, was performed. The weak pressure dependence of nitric oxide formation through the $N_2O$ pathway is strongly coupled to the termolecular reaction $H + O_2(+M) \leftrightarrow HO_2(+M)$, which gains in importance at higher pressures. This reaction, by inhibiting the formation of atomic oxygen and depleting the pool of hydrogen radicals, hampers the formation of $N_2O$ molecules, and hinders their transformation into nitric oxide. This conclusion, based on simulations, should be supported by future measurements of O, H, and $N_2O$ concentrations.

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of Canada (NSERC), the Fonds de recherche du Québec—Nature et technologies (FRQNT), and Siemens Energy Canada Ltd.

## Nomenclature

### Thermochemical Models

ARG = Argonne National Laboratory and Technical University of Denmark [25]
CRECK = Politecnico di Milano [24]
CSE = Combustion Science and Engineering, Inc. [20]
GRI = GRI-Mech 3.0 [18]
KON = Lund University [23]
NUIG = National University of Ireland, Galway [21,22]
SD = University of California, San Diego [19]

### Symbols

$A_{21}$ = rate constant of spontaneous emission, $s^{-1}$
$B_{12}$ = Einstein constant of photon absorption, $m^2J^{-1}s^{-1}$
BC = boundary condition
$C_{Bkg}$ = background coefficient
$C_{opt}$ = optical constant of the LIF system, $m^3$
$D$ = throat diameter of the inner nozzle, m
$E_L$ = mean of the time-integrated laser pulse power profiles, J
$f_B$ = Boltzmann fraction of molecules in the state excited by the laser
$f_{LIF}$ = number of photons emitted per unit molecule and laser energy, $J^{-1}$
$F_{Bkg_1}$ = laser-dependent interfering signal per unit laser energy, $J^{-1}$
$F_{NO}$ = NO-LIF signal per unit laser energy, $J^{-1}$
$k_F$ = specific rate in the forward direction, mol $m^3$ s
$k_j$ = specific rate of reaction $j$, mol $m^3$ s
$l$ = length of the computational domain, m
$L$ = nozzle-to-plate separation distance, m
L.S. = logarithmic sensitivity
$N$ = number density, $m^{-3}$
$P$ = pressure, atm
$q_B$ = backward reaction rate, mol/$m^3$ s
$q_F$ = forward reaction rate, mol/$m^3$ s
$q_{Net}$ = net reaction rate, mol/$m^3$ s
$Q_{21}$ = rate constant of non-radiative collisional quenching, $s^{-1}$
$R$ = universal gas constant, $m^3$·atm/mol-K
$S_{Bkg_0}$ = raw, laser-independent background signal
$S_{NO}$ = raw NO-LIF signal
$S_{Raw}$ = raw signal measured by the camera
$T$ = temperature, K
TBL = thermal boundary layer
$u$ = axial velocity, m/s
$X_i$ = mole fraction of species $i$
$Y_i$ = mass fraction of species $i$
$z$ = axial position, m
$\Gamma$ = dimensionless overlap fraction
$\Delta\nu_L$ = spectral width of the laser, $cm^{-1}$
$\lambda$ = laser wavelength, nm
$\rho$ = density, kg/$m^3$
$\tau$ = residence time downstream of the flame front, s
$\phi$ = equivalence ratio

### Subscripts

inlet = at the inlet of the computational domain
s = in the seeded flame
wall = at the stagnation surface

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